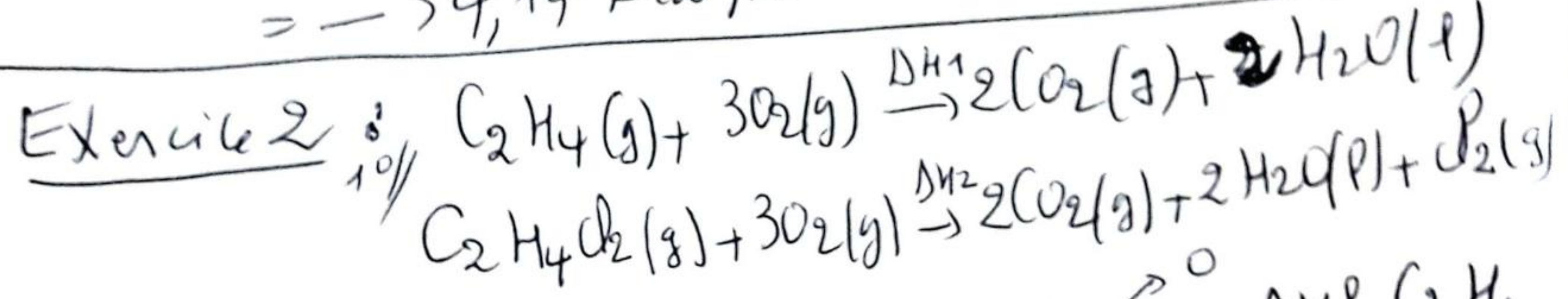


$$\Rightarrow 2\Delta H_{subl} - 3E_{H-H} + 6E_{C-H} + E_{C-C} - \Delta H_f^\circ(C_2H_6) = 0$$

$$\Rightarrow E_{C-C} = \Delta H_f^\circ(C_2H_6) - 6E_{C-H} + 3E_{H-H} - 2\Delta H_{subl}$$

$$= -24,26 - 6 \times (-99,5) + 3(-104) - 2 \times 171,7$$

$$= -34,14 \text{ Kcal/mol}$$



20// $\Delta H_1 = 2\Delta H_f^\circ H_2O + 2\Delta H_f^\circ CO_2 - 3\Delta H_f^\circ O_2 - \Delta H_f^\circ C_2H_4$

$\Delta H_2 = \Delta H_f^\circ Cl_2 + 2\Delta H_f^\circ H_2O + 2\Delta H_f^\circ CO_2 - 3\Delta H_f^\circ O_2 - \Delta H_f^\circ C_2H_4Cl_2$

par conséquent: $\Delta H_f^\circ C_2H_4 = 2(-94) + 2(-68,4) - (-331,6) = 6,8 \text{ Kcal/mol}$

$\Delta H_f^\circ C_2H_4Cl_2 = 2(-94) + 2(-68,4) - (-367,1) = 42,3 \text{ Kcal/mol}$

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